

Name: _____

Date: _____

Balancing Equations

Balance the following chemical equations.

- $\underline{2} \text{ Fe} + \underline{3} \text{ H}_2\text{SO}_4 \longrightarrow \underline{1} \text{ Fe}_2(\text{SO}_4)_3 + \underline{3} \text{ H}_2$
- $\underline{2} \text{ C}_2\text{H}_6 + \underline{7} \text{ O}_2 \longrightarrow \underline{6} \text{ H}_2\text{O} + \underline{4} \text{ CO}_2$
- $\underline{3} \text{ KOH} + \underline{1} \text{ H}_3\text{PO}_4 \longrightarrow \underline{1} \text{ K}_3\text{PO}_4 + \underline{3} \text{ H}_2\text{O}$
- $\underline{1} \text{ SnO}_2 + \underline{2} \text{ H}_2 \longrightarrow \underline{1} \text{ Sn} + \underline{2} \text{ H}_2\text{O}$
- $\underline{4} \text{ NH}_3 + \underline{5} \text{ O}_2 \longrightarrow \underline{4} \text{ NO} + \underline{6} \text{ H}_2\text{O}$
- $\underline{2} \text{ KNO}_3 + \underline{1} \text{ H}_2\text{CO}_3 \longrightarrow \underline{1} \text{ K}_2\text{CO}_3 + \underline{2} \text{ HNO}_3$
- $\underline{1} \text{ B}_2\text{Br}_6 + \underline{6} \text{ HNO}_3 \longrightarrow \underline{2} \text{ B}(\text{NO}_3)_3 + \underline{6} \text{ HBr}$
- $\underline{2} \text{ BF}_3 + \underline{3} \text{ Li}_2\text{SO}_3 \longrightarrow \underline{1} \text{ B}_2(\text{SO}_3)_3 + \underline{6} \text{ LiF}$
- $\underline{4} (\text{NH}_4)_3\text{PO}_4 + \underline{3} \text{ Pb}(\text{NO}_3)_4 \longrightarrow \underline{1} \text{ Pb}_3(\text{PO}_4)_4 + \underline{12} \text{ NH}_4\text{NO}_3$
- $\underline{1} \text{ SeCl}_6 + \underline{1} \text{ O}_2 \longrightarrow \underline{1} \text{ SeO}_2 + \underline{3} \text{ Cl}_2$