

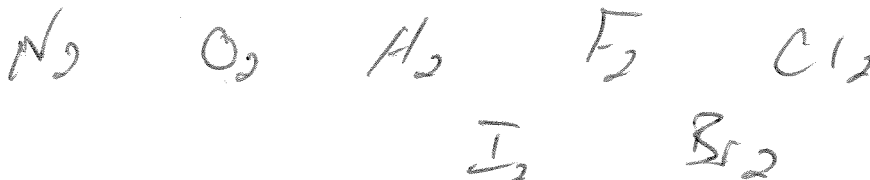
Nov. 26th/27th Warm Up Quiz

Name:

ANNE S. S

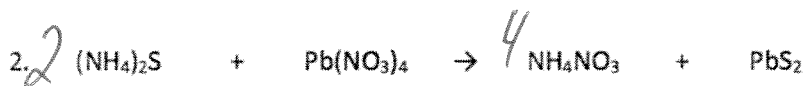
Part A:

1. What are 7 diatomic molecules?



Part B:

Balance the following equations:

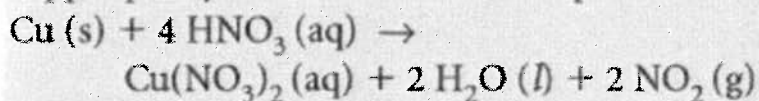


From the word statements, create the chemical formulas, write and balance the chemical equations:

3. Lead and hydrogen phosphate combine to form lead II phosphate and hydrogen gas



The following describes the reaction between a copper penny and nitric acid in an open beaker:



The mass of the beaker and contents was monitored until the reaction ended. The results are recorded in Table 2.

Table 2

Mass of beaker and contents (g)	Time (s)
224.6	0.0
215.8	118.6

- (a) What chemical is represented by the change in mass? Explain why it *cannot* be the mass change of copper.



IT IS A METAL WHICH CANNOT BE TURNED INTO A GAS!

- (b) Calculate the rate of reaction and express your answer with appropriate units.

$$\frac{8.8 \text{ g}}{118.6} = 0.07 \text{ g/s}$$

- (c) Would this reaction be considered a closed system? Why or why not?

No... BECAUSE THE GAS WAS LOST!