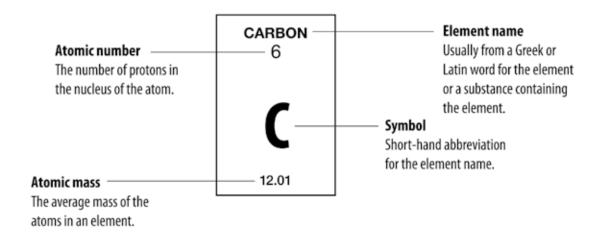
Activity Sheet
Chapter 4, Lesson 2
The Periodic Table

Name	
Date	

Your group will receive a set of cards with information that describes a particular atom. Your job is to figure out which atom the card describes and to place it in the area in your classroom for that atom.

You will use the Periodic Table, Elements 1–20 chart to help you determine what atom your card describes. The diagram and information below will help you match your cards to the correct atoms.



Parts of an Atom

Proton

Positively charged particle in the nucleus of the atom.

The number of protons in an atom's nucleus is the atomic number.

Electron

Negatively charged particle surrounding the nucleus of the atom.

The number of electrons surrounding the nucleus of an atom is equal to the number of protons in the atom's nucleus.

Neutron

Particle in the nucleus that has about the same mass as a proton but has no charge. For the atoms of the first 20 elements, the number of neutrons is either equal to or slightly greater than the number of protons.

Placing your cards

Once you know what the information in each box on your periodic table stands for and you know the parts of the atom, you will be able to correctly place most of your cards with the atoms they describe. You will need to know the following additional information in order to answer any question having to do with neutrons.

To match the number of neutrons listed on your card to the correct element, look for an element on the periodic table so that if you add the number of neutrons on your card to the protons of the element, you will get close to the atomic mass for that element.

For example, you may have a card that says, "The atom you are looking for has 5 neutrons." Look at the periodic table to find an atom that you could add 5 to its number of protons that would give you a sum close to the atomic mass given for that element. The answer is beryllium (Be), which has 4 protons and an atomic mass of 9.01.

	# of	# of	#	
PERIODIC TABLE			Write the number of protons alectrons and politrons in each alement	איווכ נווכ וומוווסכו כו סוסנטווז, כוככנוטווז, מוומ ווכמנוטווז ווו כמכון כוכוווכוונ.
HYDROGEN 1	# of Protons:	# of Electrons:	# of Neutrons:	1.01

HELIUM
2
of Protons:
of Electrons:
of Neutrons:

4.00

LITHIUM	BERYLLIUM 4	BORON	CARBON 6	NITROGEN 7	OXYGEN 8	FLUORINE 9	NEON 10	
# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	
# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	
# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	
6.94	9.01	10.81	12.01	14.01	16.00	19.00	20.18	
SODIUM 11	MAGNESIUM 12	ALUMINUM 13	SILICON 14	PHOSPHORUS 15	SULFUR 16	CHLORINE 17	ARGON 18	
# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	# of Protons:	
# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	# of Electrons:	
# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	# of Neutrons:	
22.89	24.31	26.98	28.09	30.97	32.07	35.45	39.95	
POTASSIUM 19	саLсіим 20		+ c q+	Note: Domes and the total of the contract of t			90 00	

added to the number of protons, gives a sum as close as possible to the atomic mass. element. The number of neutrons you write in this chart will be a number, that when Note: Remember that the number of neutrons is not the same for every atom of an

	0000	90.04

39.10

of Electrons:

of Electrons:

of Protons:

of Protons:

of Neutrons:

of Neutrons:

																2
The Periodic Table of the Elements	The Periodic	eriodic	dic	_	[abl	e of	the	Elen	nent	S						Helium
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Lithium				4 4	Elemer	Element Symbol Element Name	<u></u>				Boron 10.81	Carbon 12.01	Nitrogen 14.01	Oxygen 16.00	Fluorine 19.00	Neon 20.18
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Magnesium 24.31)						Aluminum 26.98	Silicon 28.09	Phosphorus 30.97	Sulfur 32.07	Chlorine 35.45	Argon 39.95
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Calcium Scandium Titanium Vanadium Chromium Manganese 40.08 44.96 47.87 50.94 52.00 54.94	Vanadium Chromium 50.94 52.00	Chromium 52.00	_	Mangi 54.	inese 94	Iron 55.85	Cobalt 58.93	Nickel 58.69	Copper 63,55	Zno 65.39	Gallium 69.72	Germanium 72.61	Arsenic 74.92	Selenium 78.96	Bromine 79.90	Krypton 83.80
39 40 41 42	41 42	42		⁴	_	44	45	46	47	48	49	20	51	52	53	54
Zr Nb Mo	QN QN	<u>§</u>		Ξ	0	χ	Кh	Pa	Ag	3	=	S	Sp	<u>e</u>	_	×
Strontlum Yttrlum Zircorilum Niobium Molybednum Technetlum 87.62 88.91 91.22 92.91 95.94 (98)	Niobium Molybednum 92.91 95.94	Molybednum 95.94		Technel (98)	mg (Rutherium 101.07	Rhodium 102.91	Palladium 106.42	Silver 107.87	Cadmium 112.41	Indium 114.82	Tin 118.71	Antimony 121.76	Tellurium 127.60	lodine 126.90	Xenon 131.29
57 72 73 74	73 74	74	Γ.	۱۳		9/	77	78	79	80	15	82	83	84	82	98
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89 104 105 106	105 106	106	+	107		108	109	110	111							,
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40.12	140.91	144.24	(145)	150.36	151.96	157.25	158.93	162.50	164.93	167.26	168.93	173.04	174.97
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E	Proactinium	Uranium	Neptunium	Plutonium	Americium	Ourium	Berkelium	Californium	Einstenium	Fermium	Mendelevium	Nobelium	Lawrencium
攻	231.04	238.03	(237)	(244)	(243)	(247)	(247)	(251)	(252)	(257)	168.93	(259)	(262)